08b Percent Yield of Carbon Dioxide Lab 2.0 (1642870)

Question 1 2 3 4 5 6

Description

A mass of potassium carbonate near 2.500 g is added to 50.0 mL of 1.00 M HCl (which equals 0.0500 moles of HCl) in a 250 mL erlenmeyer flask. Carbon dioxide gas is produced and escapes into the atmosphere.

Instructions

Use this periodic table for all calculations: Basic Periodic Table

Enter the numbers in order. Some of the data is used to calculate later numbers.

1. Question Details — Unit 8 Percent Yield Lab Partner Name [1691971]

Who was your lab partner for this lab? (enter "none" if you worked by yourself)

2. Question Details — Percent Yield Lab #2 (procedure) [3764932]

List the experimental procedure (stepwise) that should be followed in this lab (not the calculations).

3. Question Details — Unit 8 Percent Yield Lab Balance Equation [1688911]

Write the balanced equation for the reaction in this lab. (Use the lowest possible coefficients. Omit states-of-matter in your answer.)

4. Question Details — Unit 8 Percent Yield Lab Limiting reactant [1688915]

What is the limiting reactant in this experiment?

- ○ potassium carbonate
- ○ hydrochloric acid

5. Question Details — Unit 8 Percent Yield Lab Excess Reactant [1689244]

If a student weighed out 2.500 g of potassium carbonate, what is the amount of excess reactant remaining in the flask? 4.0 ✓

mol

6. Question Details Unit 8 Percent Yield Lab Data New [1691939]

a. Enter the experimental mass(from the balance) of the empty weighing dish: [ ] g

b. Enter the experimental mass(from the balance) of the weighing dish and the potassium carbonate: [ ] g

c. Enter the experimental mass of the potassium carbonate: [ ] g

Submit question c. before proceeding. The answer to question c. must be correct before the other questions can be properly evaluated.

d. Enter the number of moles of hydrochloric acid in the flask: [ ] moles

e. What is the theoretical yield of CO_2? [ ] g

Submit question c. before proceeding. The answer to question c. must be correct before the other questions can be properly evaluated.

f. Enter the mass of the flask and hydrochloric acid: [ ] g

g. Enter the mass of the flask and products(when the experiment is complete): [ ] g

h. Enter the actual yield of carbon dioxide: [ ] g

i. Enter the percent yield for this reaction: [ ] %

Assignment Details

Name (AID): **08b Percent Yield of Carbon Dioxide Lab 2.0 (1642870)**
Submissions Allowed: **7**
Category: **Lab**
Code:
Locked: **Yes**
Author: **Ryan, Matt (** mryan@allsaintsschool.org **)**
Last Saved: **Jan 25, 2018 09:26 AM CST**
Group: **Coronado High School**
Randomization: **Person**
Which graded: **Last**

Feedback Settings
Before due date
Question Score
Assignment Score
Publish Essay Scores
Question Part Score
Mark
Add Practice Button
Help/Hints
Response
Save Work
After due date
Question Score
Assignment Score
Publish Essay Scores
Key
Question Part Score
Solution
Mark
Add Practice Button
Help/Hints
Response