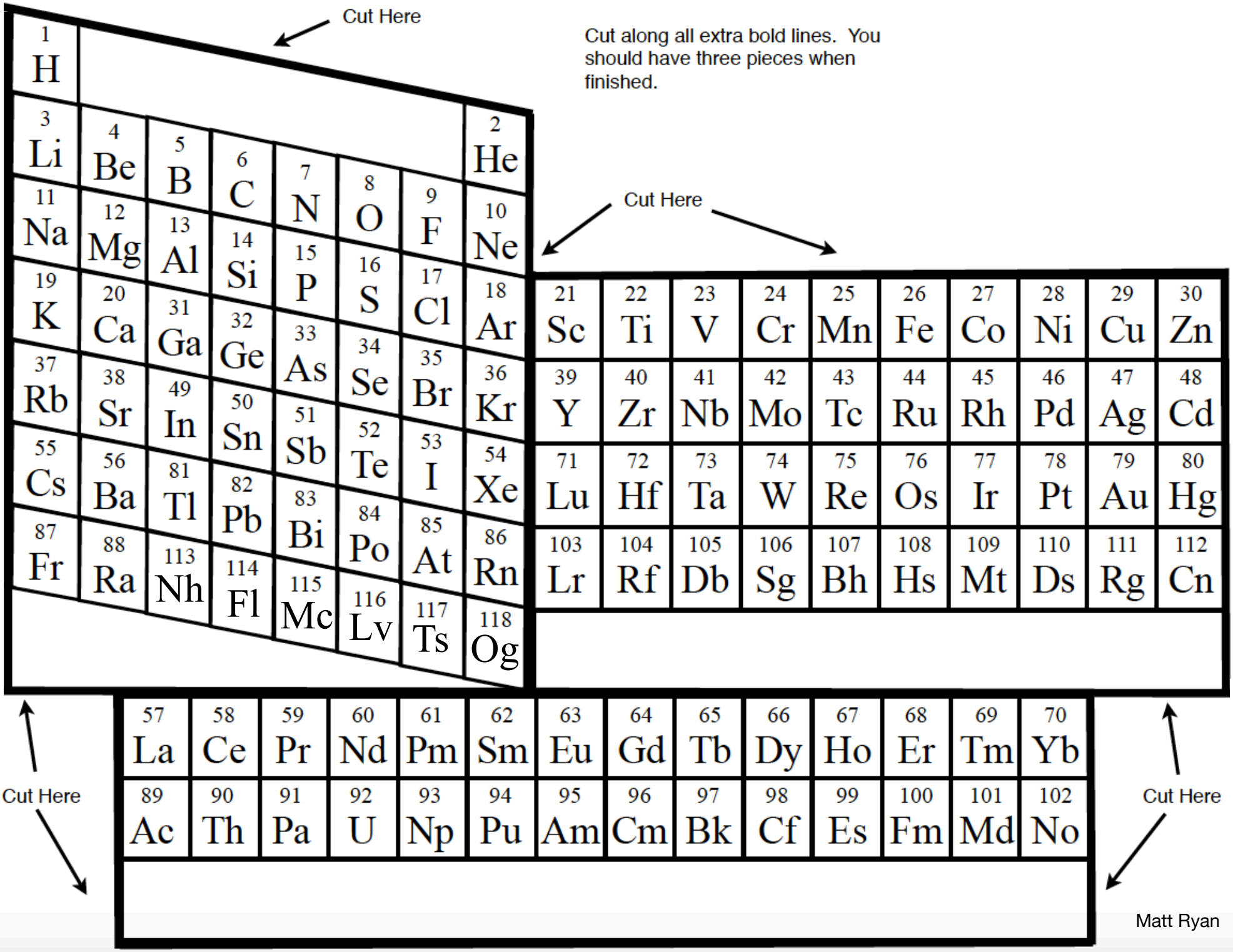
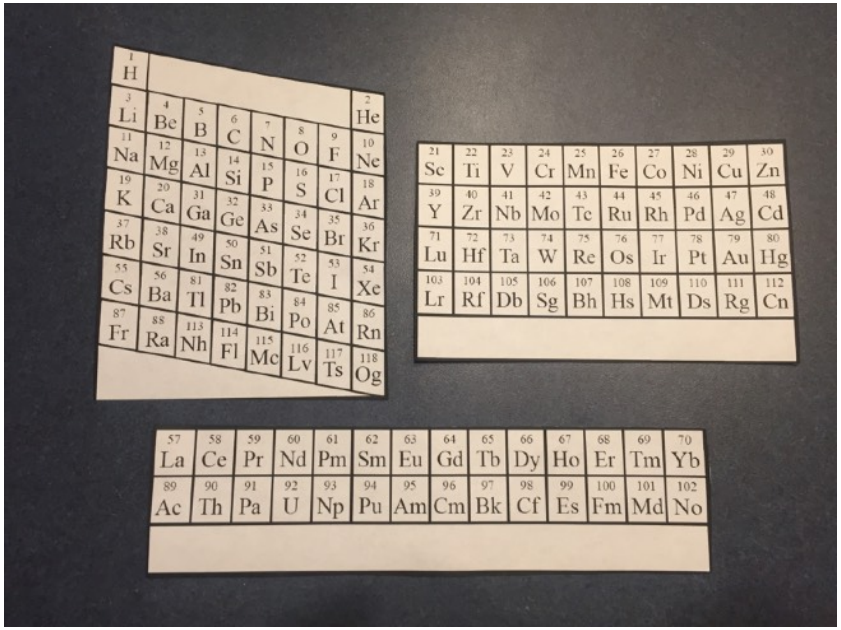
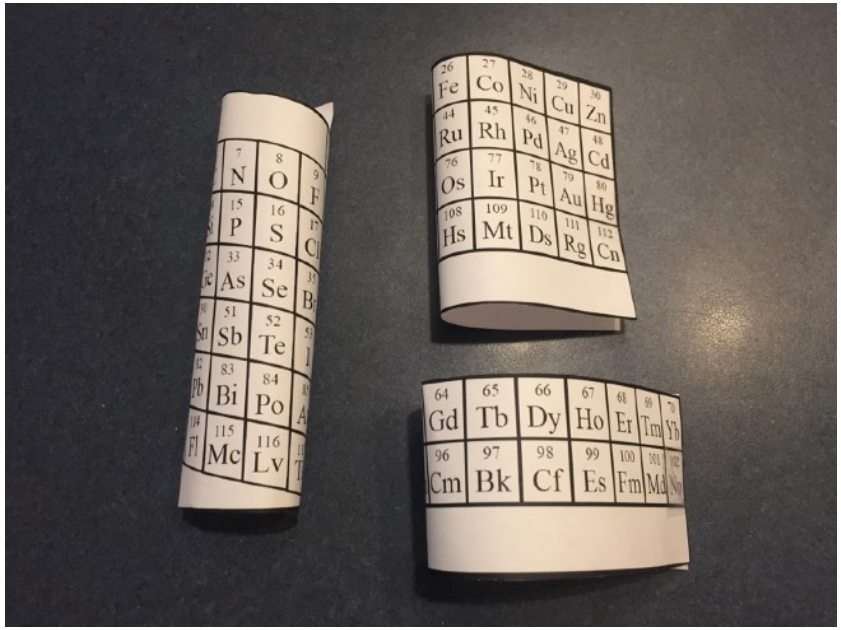
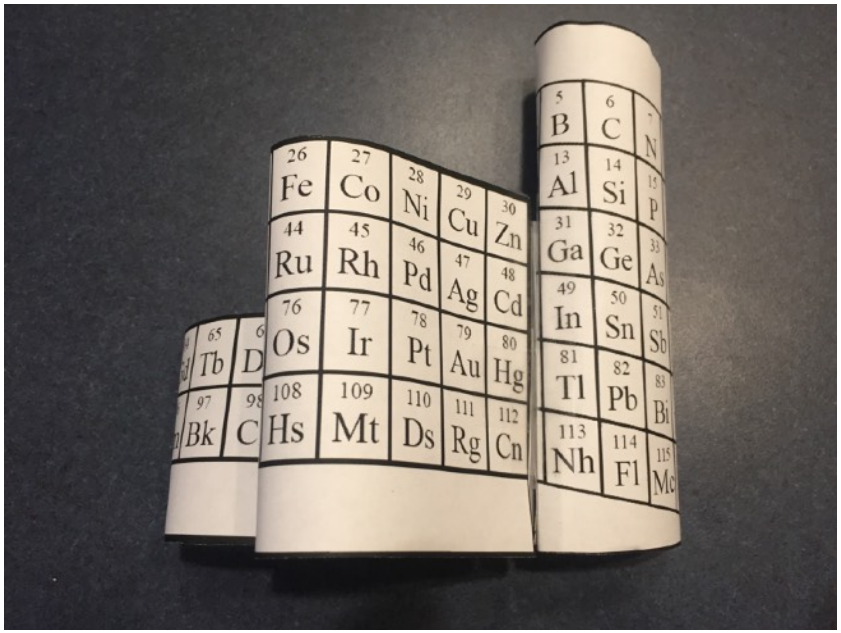

Cut Here
Cut along all extra bold lines. You should have three pieces when finished.
Cut Here
Cut Here
Cut Here
1 H
2 He
3 Li
4 Be
5 B
6 C
7 N
8 O
9 F
10 Ne
11 Na
12 Mg
13 Al
14 Si
15 P
16 S
17 Cl
18 Ar
19 K
20 Ca
31 Ga
32 Ge
33 As
34 Se
35 Br
36 Kr
37 Rb
38 Sr
49 In
50 Sn
51 Sb
52 Te
53 I
54 Xe
55 Cs
56 Ba
81 Tl
82 Pb
83 Bi
84 Po
85 At
86 Rn
87 Fr
88 Ra
113 Nh
114 Fl
115 Mc
116 Lv
117 Ts
118 Og
21 Sc
22 Ti
23 V
24 Cr
25 Mn
26 Fe
27 Co
28 Ni
29 Cu
30 Zn
39 Y
40 Zr
41 Nb
42 Mo
43 Tc
44 Ru
45 Rh
46 Pd
47 Ag
48 Cd
71 Lu
72 Hf
73 Ta
74 W
75 Re
76 Os
77 Ir
78 Pt
79 Au
80 Hg
103 Lr
104 Rf
105 Db
106 Sg
107 Bh
108 Hs
109 Mt
110 Ds
111 Rg
112 Cn
57 La
58 Ce
59 Pr
60 Nd
61 Pm
62 Sm
63 Eu
64 Gd
65 Tb
66 Dy
67 Ho
68 Er
69 Tm
70 Yb
89 Ac
90 Th
91 Pa
92 U
93 Np
94 Pu
95 Am
96 Cm
97 Bk
98 Cf
99 Es
100 Fm
101 Md
102 No
Matt Ryan

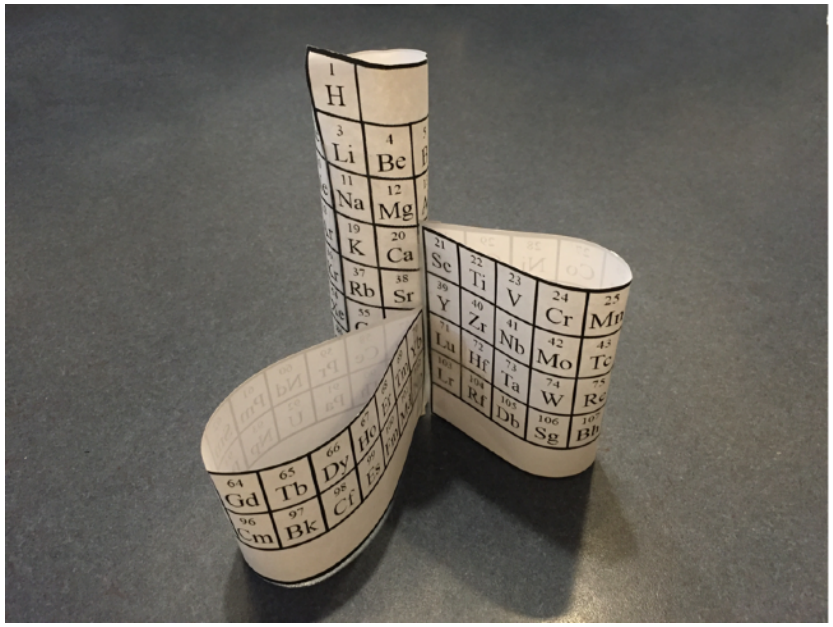
Three pieces remain after cutting along the extra bold lines.



Roll the main group elements into a tube and tape the edges such that the atomic numbers are sequential. Fold the other two pieces and tape the edges.







Teaching Suggestions:

Use this periodic table:

to allow students make their own model

to show that there are many versions of the table

to illustrate that elements such as noble gases and alkali metals are neighbors (1 electron difference)

to demonstrate a better arrangement (compare to a flat periodic table) of the transition metals, lanthanides, and actinides.